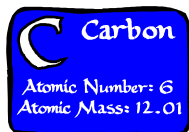


PARTS OF AN ATOM

Name _____ Per _____

An atom is made up of protons and neutrons which are in the nucleus, and electrons which are in the electron cloud surrounding the atom. The atomic number equals the number of protons. The electrons in a neutral atom equal the number of protons. The mass number equals the sum of the protons and neutrons. The charge indicates the number of electrons that have been lost or gained. A positive charge indicates the number of electrons (which are negatively charged) lost. A negative charge indicated the number of electrons gained.



This carbon atom would have 6 protons, 6 neutrons, and 6 electrons with a charge of 0.

Complete the following chart.

Element/ion	Atomic Number	Mass Number	Charge	Protons	Neutrons	Electrons
Mg	12	24	0	12	12	12
K						
Na ⁺¹						
F ⁻¹						
Al ⁺³						
H						
Mg ⁺²						
Ag						
S ⁻²						
Cu						
Cl ⁻¹						
Be ⁺²						

Write the symbols for the following elements.

1. oxygen _____
2. hydrogen _____
3. chlorine _____
4. sodium _____
5. fluorine _____
6. carbon _____
7. helium _____
8. nitrogen _____
9. copper _____
10. sulfur _____

11. magnesium _____
12. manganese _____
13. neon _____
14. bromine _____
15. phosphorus _____
16. silver _____
17. lead _____
18. iron _____
19. calcium _____
20. potassium _____

Write the name of the element that corresponds to each of the following symbols.

21. Cu _____
22. K _____
23. C _____
24. Au _____
25. Zn _____
26. Pb _____
27. Fe _____
28. Na _____
29. S _____
30. Al _____

31. Ca _____
32. Ag _____
33. P _____
34. O _____
35. I _____
36. Sn _____
37. H _____
38. F _____
39. Ni _____
40. Hg _____